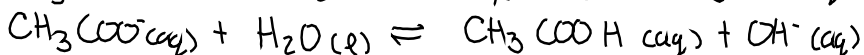
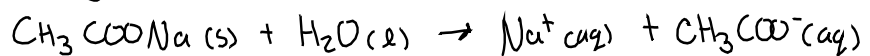


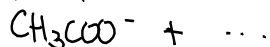
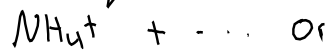
- Salts that yield basic solns

- Salt of a weak acid & a strong base yields a basic soln in H_2O because the anion acts as a weak base & the cation does not react

- ex: CH_3COONa is basic



- This is why, when you see a problem that has you working with, say, NH_4Cl or CH_3COONa you can write the equilibrium equation as



immediately w/o worrying about dealing w/ the Na^+ or Cl^- ions

Problems:

Acidic, basic, or neutral

1. $KClO_4$
2. C_6H_5COONa
3. $CrCl_3$
4. $KClO_2$
5. $CH_3NH_3NO_3$
6. CsI

Solns:

1. Neutral $\rightarrow K^+$ & ClO_4^- are ions from strong base & acid
2. Basic $\rightarrow C_6H_5COONa \rightleftharpoons C_6H_5COO^- + H_2O \rightleftharpoons C_6H_5COOH + OH^-$
3. Acidic $\rightarrow Cr^{3+}$ is small ion w/ high charge $\Rightarrow Cr(H_2O)_6^{3+} + H_2O \dots$
4. Basic $\rightarrow ClO_2^- + H_2O \rightleftharpoons HClO_2 + OH^-$, K^+ from strong base
5. Acidic $\rightarrow CH_3NH_3^+ + H_2O \rightleftharpoons CH_3NH_2 + H_3O^+$, NO_3^- from strong acid
6. Neutral $\rightarrow Cs^+$ from strong base, I^- from strong acid

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