

Anions of Strong acids:

- halide ions (except F^-)

Anions of Strong oxo acids:

- NO_3^- , ClO_4^- , HSO_4^-

Cations of Strong bases:

- group 1A(1) + group 2A(2) (except Be^{2+})

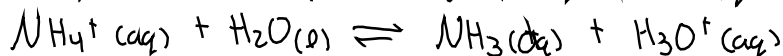
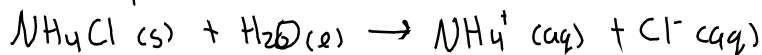
Salts containing only these ions yield neutral solns because no rxn w/ H_2O takes place

- ex: $NaCl$, $Ba(NO_3)_2$

- Salts that yield acidic solns

- Salt of a strong acid + a weak base yields an acidic soln because the cation acts as a weak acid + the anion does not react

- ex: NH_4Cl is acidic



- Cl^- is anion of strong acid \rightarrow neutral

- NH_4^+ is a weak acid (conj acid of NH_3 weak base)

- Small, highly charged metal cations

- ex: $Fe(NO_3)_3$

$Fe^{3+} \rightarrow$ weak acid

$NO_3^- \rightarrow$ anion of strong acid

(See prev. page Al^{3+} example, this does the same thing)